

SYLLABUS

1. Information regarding the programme

1.1 Higher education institution	Babes-Bolyai University
1.2 Faculty	Chemistry and Chemical Engineering
1.3 Department	Chemical Engineering
1.4 Field of study	Chemical Engineering
1.5 Study cycle	Master
1.6 Study programme / Qualification	Advanced Chemical Process Engineering

2. Information regarding the discipline

2.1 Name of the discipline	Advanced Physical Chemistry – CME6111						
2.2 Course coordinator	Assoc. Prof. dr. eng. Adrian Nicoară (Thermodynamics „TD” and Electrochemistry) Assoc. Prof. Dr. Eng. Alexandra Csavdări (Chemical Kinetics “CK”)						
2.3 Seminar / laboratory work coordinator	Assoc. Prof. dr. eng. Adrian Nicoară (Thermodynamics „TD” and Electrochemistry) Assoc. Prof. Dr. Eng. Alexandra Csavdări (Chemical Kinetics “CK”)						
2.4. Year of study	I	2.5 Semester	1	2.6. Type of evaluation	E	2.7 Type of discipline	Ob

3. Total estimated time (hours/semester of didactic activities)

3.1 Hours per week	2	Of which: 3.2 course	2	3.3 seminar/laboratory	2/0
3.4 Total hours in the curriculum	56	Of which: 3.5 course	28	3.6 seminar/laboratory	28/0
Time allotment:					hours
Learning using manual, course support, bibliography, course notes					31
Additional documentation (in libraries, on electronic platforms, field documentation)					7
Preparation for seminars/labs, homework, papers, portfolios and essays					14
Tutorship					14
Evaluations					3
Other activities: not the case					-
3.7 Total individual study hours	69				
3.8 Total hours per semester	125				
3.9 Number of ECTS credits	5				

4. Prerequisites (if necessary)

4.1. curriculum	<ul style="list-style-type: none"> Not the case
4.2. competencies	<ul style="list-style-type: none"> Not the case

5. Conditions (if necessary)

5.1. for the course	<ul style="list-style-type: none"> • The students will turn off their mobile phones or put them on silent mode • Delays will not be tolerated
5.2. for the seminar /lab activities	<ul style="list-style-type: none"> • Students will bring to the laboratory practice their course notes and appropriate calculus device • Students will turn off their mobile phones or put them on silent mode • Delays will not be tolerated

6. Specific competencies acquired

Professional competencies	<ul style="list-style-type: none"> • Definition of notions, concepts, theories and advanced models in the field of chemistry and chemical process engineering as well as their adequate use within the professional community. • Use of advanced knowledge in the field of chemistry and chemical process engineering to explain and interpret chemical processes. • Identification and proper usage of concepts, method and theories for solving new complex problems of chemical process engineering. • Critical analysis and usage of principles, methods and advanced work techniques for quantitative and qualitative evaluation of chemical process engineering.
Transversal competencies	<ul style="list-style-type: none"> • Independent execution of complex professional duties and research projects using computer-aided techniques and comply with professional ethics and moral. • Planning, monitoring and assuming professional duties of underline group. Proving the coordination capabilities, analytical thinking, adaptability and flexibility, collaboration with team members. • Auto-evaluation of professional performances and establish the needs of continuous learning, documentation in the work fields in correlation with the labour market.

7. Objectives of the discipline (outcome of the acquired competencies)

7.1 General objective of the discipline	<ul style="list-style-type: none"> • Approach of advanced concepts of physical chemistry (thermodynamics, chemical kinetics, electrochemistry).
7.2 Specific objective of the discipline	<ul style="list-style-type: none"> • Advanced approach in the field of thermodynamics of reversible and of irreversible processes. • Correlation of advanced thermodynamics fundamentals and the ability to use / apply / correlate theoretical knowledge and interpret phenomena and processes associated in the field. • Approach of advanced concepts of electrochemistry: the electric double layer; electro-capillary and electro-kinetic phenomena; types of overpotential; reactions under mixed control (activation + diffusion). • Training of students to use electrochemical investigation methods for the electrode processes; Cyclic voltammetry. • Interpretation of kinetic data from the point of view of rate laws and reaction mechanisms. • Interpretation of kinetic data for complex reaction systems in homogeneous and heterogeneous media.

8. Content

8.1 Course	Teaching methods	Remarks
8.1.1. TD1: Review of the main consequences of first and second laws of thermodynamics.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.2. TD2: Introduction in irreversible processes thermodynamics. Basic concepts. Local equilibrium hypostesis. Entropy balance.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.3. TD3: General theory of irreversible process thermodynamics: 7 steps formulation.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.4. TD4: Minimum entropy production principle. Application to heat transport.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.5. TD5: Irreversible thermodynamics of electrokinetic phenomena and membrane processes.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.6. EC1: Recap of fundamental concepts in electrochemistry. The electric double layer. Double layer and electrokinetic phenomena.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.7. EC2: Advanced electrochemical kinetics. Types of overpotential. Electrode kinetic law for multielectrone transfer processes.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.8. EC3: Horiuti number. Mass transport overpotential.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.9. EC4: Electrochemical investigation methods of electrode processes (classification, examples) and cyclic voltammetry.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.10. CK1: Recap of fundamental concepts in chemical kinetics and the bond to applications in chemical engineering.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.11. CK2: Methods of obtaining experimental data Processing kinetic experimental data – general concepts, determination of reaction orders and rate coefficients. Empirical rate laws.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.12. CK3: Empirical rate laws – continuation. Interpretation of rate laws from the point of view of reaction mechanisms.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.13. CK4: Determination of individual rate coefficient for some complex reaction systems in homogeneous media.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours
8.1.14. CK5: Determination of individual rate coefficient for some complex reaction systems in heterogeneous media.	Presentation; Explanation; Conversation; Description; Debate	Alocated time = 2 hours

Bibliography

1. A. Kalyan, I.K. Puri, "Advanced thermodynamics Engineering", CRC Press, 2002.
2. P.W. Atkins, "Physical Chemistry", any Edition
3. I.G. Murgulescu, R. Valcu, "Introducere in chimia fizică. Termodinamica chimica", vol III, Ed. Academiei RSR, Bucuresti, 1982.
4. I. Baldea, „Deducerea mecanismului de reacție”, Presa Universitară Clujeană, Cluj-Napoca, 2008.
5. I. Baldea, „Cinetica chimica si mecanisme de reactie. Baze teoretice si aplicatii”, Presa Universitara Clujeana, Cluj-Napoca, 2002.

6. I. Baldea, „Some advanced topics in chemical kinetics”, Cluj University Press, 2000.
7. G. Bozga, O. Muntean, „Reactoare chimice”, Vol. I + II, Editura Tehnică, București, 2006.
8. O. Levenspiel, “Chemical Reactor Engineering”, Third Edition, John Wiley & Sons, 1999.
9. L. Oniciu, E. Constantinescu, „Electrochimie si coroziune”, Editura Didactica si Pedagogica, Bucuresti, 1987.
10. L. Oniciu, L. Muresan, „Electrochimie aplicata”, Presa Universitară Clujeană, Cluj-Napoca, 1998.
11. PPT presentations – available in revised form during the semester

8.2 Seminar	Teaching methods	Remarks
8.2.1. TD1: Applications to first law of thermodynamics.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.2. TD2: Applications to second law of thermodynamics.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.3. TD3: Applications to general irreversible thermodynamics.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.3. TD4: Numerical applications of irreversible processes thermodynamics: heat transfer examples.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.4. TD5: Numerical applications of irreversible processes thermodynamics: membrane processes examples.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.6. EC1: Double layer numerical examples.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.7. EC2: Electrode kinetics applications: activation control.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.8. EC3: Electrode kinetics applications: mass transport control.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.9. EC4: Investigation methods.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.10. CK1: Determination of partial reaction orders and individual rate coefficients from diverse kinetic experimental data, by using adequate linearization or complex kinetic models. Interpretation of experimental rate laws from the point of view of reaction mechanism – part I.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.11. CK2: Determination of partial reaction orders and individual rate coefficients from diverse kinetic experimental data, by using adequate linearization or complex kinetic models. Interpretation of experimental rate laws from the point of view of reaction mechanism – part II.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.12. CK3: Case studies – calculus of individual rate coefficients for various reaction schemes in homogeneous media.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.13. CK4: Case studies – calculus of individual rate coefficients for various reaction schemes in heterogeneous media.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours
8.2.14. Recap and preparation for the exam.	Explanation, Conversation; Description; Debate; Problem solving	Alocated time = 2 hours

Bibliography

1. P.W. Atkins, P.W. Atkins, "Physical Chemistry", any Edition
2. H. E. Avery, D. J. Shaw, "Basic Physical Chemistry Calculations", Butterworth & Co., 1980.
3. I. Baldea, „Cinetica Chimica si mecanisme de reactie. Baze teoretice si aplicatii”, Presa Universitara Clujeana, Cluj-Napoca, 2002.
4. I. Baldea, „Some advanced topics in chemical kinetics”, Cluj University Press, 2000.
5. G. Niac, V. Voiculescu, I. Baldea, M. Preda, „Formule tabele probleme de chimie fizică”, Editura Dacia, Cluj-Napoca, 1984.
6. Notes provided by lecturers – available during the semester.

9. Corroborating the content of the discipline with the expectations of the epistemic community, professional associations and representative employers within the field of the program

- By instructing the theoretical and practical concepts of **Advanced Physical Chemistry** course, the students will get the knowledge in accordance with the competencies of the Diploma Supplement and the qualifications of the ANC.

10. Evaluation

Type of activity	10.1 Evaluation criteria	10.2 Evaluation methods	10.3 Share in the grade (%)
10.4 Course	Correctness of answers – proper understanding and learning of concepts discussed during lectures; Correct use of learned concept within new contexts.	Written colloquia consisting of three sets of subjects corresponding to the TC, CC and EC modules of the course.	100 % (Each module contributes to the final mark with one third that is with 33.3 %)
	Correct solving of problems as inherent part of examination subjects.	Proven or intended fraud is punished according to the ECST rules of UBB.	
10.5 Laboratory Practice	Correctness of answers – proper understanding and learning of concepts discussed during class; Correct use of learned concept within new contexts.	Evaluated by means of problems to be solved, as inherent part of the examination subjects.	-

Minimum performance standards
<ul style="list-style-type: none">• Grade 5 (five) at the written exam, at each of the three separate modules of the course (TD, EC and CK). The final mark represents the rounded value of the average of the marks obtained at each of the three modules.• Adequate knowledge and usage of basic concepts of advanced physical chemistry.

Date

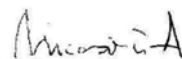
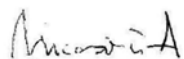
Signature of course coordinator

Signature of seminar coordinator

April the 15th, 2020

Assoc. Prof. Dr. Eng. Adrian Nicoară

Assoc. Prof. Dr. Eng. Adrian Nicoară



Assoc. Prof. Dr. Eng. Alexandra Csavdári

Assoc Prof. Dr. Eng. A. Csavdári



Date of approval

Signature of the head of department

April the 16th, 2020

Prof. Dr. Eng. G. L. Turdean

